Date
Page

Electronic transitions :-

:- when UV- visible light of desired wavelength passed through the sample, a part of radiations are absorbed and cause electronic transitions

:- It e⁻s are promoted from bonding to an anti-bonding orbit

(1) σ - σ* transitions :-

:- These are high energy transitions as σ - bonds are very strong.

:- They occurs in Saturated hydrocarbons.
eg :- methane, propane etc.

(2) n - σ* transitions :-

:- They need less energy than σ - σ* transitions.

:- They occurs in Saturated molecules containg hetero atom..
eg :- Saturated alcohol, aldehyde .etc.

(3) π - π* - transitions :-

:- They occurs in unsaturated compounds eg
eg :- alkenes, alkynes, ago etc.

(4) n - π* Transitions:

- These are least energy transitions.
- They occurs in saturated aldehydes.